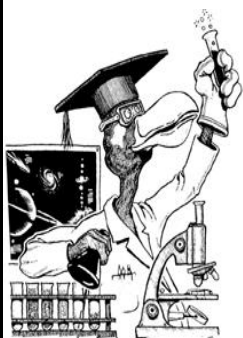


# INVESTIGATING CHEMISTRY: NAMING COMPOUNDS

NAME \_\_\_\_\_

	Formula of Compound	Ratio of elements	Name of Compound
1	CS <sub>2</sub>		
2	H <sub>2</sub> CO <sub>3</sub>		
3	SO <sub>2</sub>		
4	CF <sub>4</sub>		
5	NF <sub>3</sub>		
6	H <sub>2</sub> S		
7	CaO		
8	CaCl <sub>2</sub>		
9	CO <sub>2</sub>		
10	H <sub>2</sub> SO <sub>4</sub>		
11	Mg(OH) <sub>2</sub>		
12	CuCl <sub>2</sub>		
13	KOH		
14	CaSO <sub>4</sub>		
15	HCl		
16	Ca(OH) <sub>2</sub>		
17	NaOH		
18	AgNO <sub>3</sub>		
19	ZnCl <sub>2</sub>		
20	Cu(NO <sub>3</sub> ) <sub>2</sub>		
21	Fe <sub>2</sub> O <sub>3</sub>		
22	CuClO <sub>3</sub>		
23	Na <sub>2</sub> CO <sub>3</sub>		
24	FeCl <sub>3</sub>		

## Hints:



1. If the first element is a metal do / do not use prefixes like 'mono' and 'di'.
2. If the first element is a non-metal do / do not use prefixes like 'mono' and 'di' and 'tri'.
3. If there are only two elements in the compound the name will end in '\_\_\_\_'.
4. If there are more than two elements present in the compound, and one of them is oxygen then the name will end in '\_\_\_\_' or '\_\_\_\_'.
5. Some elements commonly form groups. Some common groups are listed below:

Combining group (radical)	Name of group
-CO <sub>3</sub>	
-NO <sub>3</sub>	
-SO <sub>4</sub>	
-PO <sub>4</sub>	
-OH	