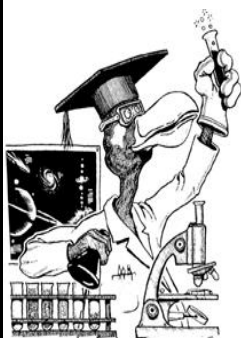


# INVESTIGATING CHEMISTRY: NAMING COMPOUNDS

NAME        **SOLUTIONS!**

	Formula of Compound	Ratio of elements	Name of Compound
1	CS <sub>2</sub>	1 carbon: 2 sulfur	Carbon disulfide
2	H <sub>2</sub> CO <sub>3</sub>	2 hydrogen: 1 carbon: 3 oxygen	Hydrogen carbonate
3	SO <sub>2</sub>	1 sulfur: 2 oxygen	Sulphur dioxide
4	CF <sub>4</sub>	1 carbon: 4 fluorine	Carbon tetrafluoride
5	NF <sub>3</sub>	1 nitrogen: 3 fluorine	Nitrogen trifluoride
6	H <sub>2</sub> S	2 hydrogen:1 sulfur	Hydrogen sulphide
7	CaO	1 calcium:1 oxygen	Calcium oxide
8	CaCl <sub>2</sub>	1 calcium:2 chlorine	Calcium chloride
9	CO <sub>2</sub>	1 carbon:2 oxygen	Carbon dioxide
10	H <sub>2</sub> SO <sub>4</sub>	2 hydrogen:1 sulfur: 4 oxygen	Sulphuric acid
11	Mg(OH) <sub>2</sub>	1 magnesium: 2 oxygen: 2 hydrogen	Magnesium hydroxide
12	CuCl <sub>2</sub>	1 copper: 2 chlorine	Copper chloride
13	KOH	1 potassium:1 oxygen:1 hydrogen	Potassium hydroxide
14	CaSO <sub>4</sub>	1 calcium: 1 sulfur:4 oxygen	Calcium sulphate
15	HCl	1 hydrogen:1 chlorine	Hydrochloric acid
16	Ca(OH) <sub>2</sub>	1 calcium:2 oxygen:2 hydrogen	Calcium hydroxide
17	NaOH	1 sodium:1 oxygen:1 hydrogen	Sodium hydroxide
18	AgNO <sub>3</sub>	1 silver:1 nitrogen:3 oxygen	Silver nitrate
19	ZnCl <sub>2</sub>	1 zinc:2 chlorine	Zinc chloride
20	Cu(NO <sub>3</sub> ) <sub>2</sub>	1 copper:2 nitrogen:6 oxygen	Copper nitrate
21	Fe <sub>2</sub> O <sub>3</sub>	2 iron:3 oxygen	Iron oxide
22	CuClO <sub>3</sub>	1 copper:1 chlorine:3 oxygen	Copper chlorate
23	Na <sub>2</sub> CO <sub>3</sub>	2 sodium:1 carbon:3 oxygen	Sodium carbonate
24	FeCl <sub>3</sub>	1 iron:3 chlorine	Iron chloride

## Hints:



1. If the first element is a metal do not use prefixes like 'mono' and 'di'.
2. If the first element is a non-metal do use prefixes like 'mono' and 'di' and 'tri'.
3. If there are only two elements in the compound the name will end in 'ide'
4. If there are more than two elements present in the compound, and one of them is oxygen then the name will end in 'ate' or 'ide'.
5. Some elements commonly form groups. Some common groups are listed below:

Combining group (radical)	Name of group
-CO <sub>3</sub>	Carbonate
-NO <sub>3</sub>	Nitrate
-SO <sub>4</sub>	Sulphate
-PO <sub>4</sub>	Phosphate
-OH	hydroxide