

# Elements, Mixtures and Compounds

MIXTURES	COMPOUNDS
1. The parts of the mixture may be present in different <u>proportion</u> .	1. In a compound, the parts are present in a <u>constant</u> proportion by weight.
2. When a mixture is made, there is no evidence of any <u>chemical</u> action taking place	2. When a compound is made there is usually evidence of chemical action e.g. <u>change in colour</u>
3. The parts of a mixture can be easily <u>separated</u> e.g. by using a magnet.	3. The parts of a compound are <u>difficult</u> to separate.
4. A mixture cannot be represented by a <u>formula</u> e.g. air or petrol	4. Always have a definite composition and <u>can</u> be represented by a formula.

## POINTS TO REMEMBER

1. An element is a substance which cannot be broken down by ordinary chemical reactions into simpler substances.
2. The smallest particles of elements are called atoms
3. Elements can be classified as metal or non-metal
4. A compound is a substance which consists of two or more elements chemically joined.
5. Every sample of a particular compound always has the same chemical composition.
6. Every compound has a formula which shows the relative numbers of each element/atom it contains.
7. When elements join to form compounds the new substance made has different properties to the original elements.

